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- Ionic Bonds and Ionic
Compounds. Ionic
Bonding. Anions and
cations are held
together by . opposite.
charges (+ and -) Ionic
compounds are called .
salts. Formula unit -
Simplest . ratio. of
elements in an ionic

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compound.

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7. Metallic atoms tend
to valence electrons to
produce a positively ...

7.2 Ionic Bonds and
Ionic Compounds

Essential

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Understanding Ionic compounds are the result of ionic bonds forming between oppositely charged ions.

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Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

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magnesium and
chlorine react, they
form an ionic
compound. What is the
charge of the ionic
compound? answer

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Metallic Bonds and
Metallic Properties

When a metal is
subjected to pressure,
the metal cations
easily slide past one
another. If an ionic
crystal is struck with a
hammer, the blow
tends to push the

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positive ions close together.

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207: 17 Answer When
it is dissolved in an
aqueous solution, the
ions are free to flow
throughout the solution
to their electrode,
allowing electric flow.

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