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## **Stoichiometry Limiting Reagent**

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**Worksheet Answer**

Step #4 Using the limiting reagent find the moles of I<sub>2</sub>

produced  $5 \text{ CO} = 1 \text{ I}_2$   
 $1.0 \text{ mol CO} \times \frac{1}{5} = 0.20 \text{ mol}$   
of I<sub>2</sub> are produced

Step #5 Find the grams of I<sub>2</sub> produced  
 $m = n \cdot M = 0.20 \text{ mol}$   
 $\cdot 253.80 \text{ g/mol} = 50.76$

grams of I<sub>2</sub> are produced Using CO as the limiting reagent, a reaction of 28.0 grams of CO will produce 50.76 grams of iodine.

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**Stoichiometric  
Worksheet #3:  
Limiting Reagents  
and ...**

Limiting Reagent

Worksheet #1 1. Given  
the following reaction:  
(Balance the equation  
first!)  $C_3H_8 + O_2$

$2 \rightarrow CO_2 + H_2O$  a)

If you start with 14.8 g  
of  $C_3H_8$  and 3.44 g of  
 $O_2$ , determine the  
limiting reagent b)  
determine the number  
of moles of carbon

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dioxide produced c)  
determine the number  
of grams of H<sub>2</sub>O  
produced

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Limiting Reagents and  
Percentage Yield

Worksheet: 1. Consider  
the reaction  $I_2O_5(g) + 5 CO(g) \rightarrow 5 CO_2(g) + I_2(g)$ : a) 80.0  
grams of iodine(V)  
oxide,  $I_2O_5$ , reacts  
with 28.0 grams of



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carbon monoxide, CO.  
Determine the mass of  
iodine I<sub>2</sub>, which could  
be produced?: b) If, in  
the above situation,  
only 0.160 moles, of  
iodine, I<sub>2</sub> was  
produced.

## **Stoichiometric Worksheet #3: Limiting Reagents and ...**

Limiting reactant  
worksheet  
stoichiometry 6 answer  
key. Which reactant is

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the limiting reagent. 7  
50 g  $\text{Na}_3\text{PO}_4$  6b. Is the  
limiting reactant  
because you have  
fewer moles of b than  
a. Download ebook  
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**Stoichiometry 6**

**Answer Key ...**

Worksheet search  
result by word Density  
worksheet with  
answers from  
stoichiometry limiting  
reagent worksheet  
answers ,  
source:ftxs8.com You  
have all your materials.  
An paper isn't unusual  
in businesses when  
they 're trying to get a  
remedy to a issue and  
will need to receive all

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the feasible  
perspectives and data  
available.  
Worksheet  
Answer Key

**Stoichiometry  
Limiting Reagent  
Worksheet Answers**

Stoichiometry -  
Limiting and Excess  
Reactant Introduction  
to Limiting Reactant  
and Excess Reactant  
The limiting reactant or  
limiting reagent is the  
first reactant to get  
used up in a chemical  
reaction. Once the

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limiting reactant gets used up, the reaction has to stop and cannot continue and there is extra of the other reactants left over.

## **Stoichiometry - Limiting and Excess Reactant (solutions**

...

Limiting Reagent  
Worksheet Answer Key  
with Work Along with  
Honors Chemistry. With  
the help of this  
question key, you

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could search for any question or any topic that you might want to explore further about.

## **Limiting Reagent Worksheet Answer Key with Work**

Limiting Reagent  
Worksheet -KEY. All of the questions on this worksheet involve the following reaction:  
When copper ... Since 13.0 grams is the smaller number, that's our answer. Please

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note that rounding differences may cause your answers to be slightly different, ... By doing a stoichiometry calculation to determine the amount of lead ...

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...

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Worksheet

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of "Limiting Reactant  
Worksheet Answers"

## **Limiting Reactant Worksheet Answers**

Simple stoichiometry  
only (one given, one  
wanted) Limiting  
reagents only (two  
given reactants, one  
wanted product) Mix &  
match (both simple  
stoichiometry and



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limiting reagent  
problems) Units to use  
(select at least one):  
Grams Moles Particles  
(e.g. atoms/molecules/  
formula units) Chemical  
formulas or names:  
Formulas only Names  
only

## **Stoichiometry & Limiting Reagents Practice Quiz | Mr ...**

Stoichiometry Limiting  
Reagent Worksheet  
Answers Limiting  
Reagents and

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Percentage Yield

Worksheet - Answers.

1. a)  $I_2O_5 + 5 CO \rightarrow$   
 $5 CO_2 + I_2$ . 80.0 g 28.0

g. Solution steps. Step #1 Determine the moles of  $I_2O_5$ . Step #2 Determine the moles of  $CO$ . Step #3 Do a Limiting Reagent Test. Step #4 Using the limiting reagent find the moles of  $I_2$  produced.

**Stoichiometry**  
**Limiting Reagent**

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## ...Worksheet

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Correctly phrased, the answer is 57 formula units. Comment: when I was in the classroom, teaching the technique for determining the limiting reagent, I would warn against using the results of the division, in this case

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the 19 for the NaOH, in the next step of the calculation. The 19 is good only for determining the limiting reagent.

## **Stoichiometry: Limiting Reagent Problems #1 - 10**

The Results for Pogil  
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Answer Key Testname:  
STOICHIOMETRY  
WORKSHEET CHEM 108

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SP 2016 1) 1.5 2) a. ...

This is a limiting reagent problem. First figure out which reagent is limiting.

$$\begin{aligned} & 39.6 \text{ g C}_2\text{H}_5\text{OH} \times \frac{1 \text{ mol}}{46.0 \text{ g}} \times 2 \text{ mol CO}_2 / 1 \text{ mol C}_2\text{H}_5\text{OH} = \\ & 1.72 \text{ mol CO}_2 \end{aligned}$$

## **Stoichiometry Worksheet Chem 30A FALL 2016**

**Name**

Worksheet 14 3

Answers to Worksheet

#14 Limiting Reagents

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## Stoichiometry

### Limiting Reagent

A Limiting Reagent is the reactant that is completely used up in a reaction. This reagent is the one that determines the amount of product formed.

Limiting reagent calculations are performed in the same manner as the stoichiometric equations on Worksheet #11. However, with a limiting



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Naming Ionic and  
Molecular Compounds |  
How to Pass

**Stoichiometry  
Worksheets Answers**

Answer \_\_\_\_ 4] For the following equation determine which reactant is the limiting reactant and which reactant is in excess. The amounts of reagent used are shown. Show

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calculations to support  
your choices.  $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$   
4 + 4H. 2. 40.0 g 16.0g  
. 40.0g Fe X . 1molFe  
55.8g X 1mol Fe<sub>3</sub>O<sub>4</sub>  
3molFe = 0.239 mol  
Fe<sub>3</sub>O<sub>4</sub>

## **WORKSHEET 13**

**Name - Cerritos**

**College**

Stoichiometry is the  
chemical term to  
describe calculations  
that allow us to find  
the amounts of

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chemicals involved in a given reaction. After you finish this worksheet, bring it to your teacher to check your answers, when finished you may make your S'more. In stoichiometry, you must always start with a balanced equation!

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Some of the  
*Page 28/30*

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worksheets for this concept are  
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Stoichiometry practice work, Balancing equations and simple stoichiometry key, Solution stoichiometry work, Chemistry computing formula mass work, Work stoichiometry and chemical formula calculations,  
Stoichiometry work 1 answers, Chapter 3

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stoichiometry.  $2 \text{SO}_4^{2-}$   
 $\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$ . 6 mol  
of aluminum reacts  
with 4.

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